

24 THIN LAYER CHROMATOGRAPHY	Page 1 of 2
Division of Forensic Science  TRACE EVIDENCE TRAINING MANUAL	Amendment Designator:
	Effective Date: 29-March-2004
<p style="text-align: center;"><b>24 THIN LAYER CHROMATOGRAPHY (TLC)</b></p> <p><b>24.1 Introduction to Thin Layer Chromatography</b></p> <p>24.1.1 Objectives</p> <p>Through completion of this module the trainee will have developed and demonstrated theoretical knowledge and/or practical skills to:</p> <ul style="list-style-type: none"> <li>• Describe the theory and basic principles of thin layer chromatography;</li> <li>• Select an appropriate mobile phase;</li> <li>• Select an appropriate detection and/or development method; and</li> <li>• Perform qualitative separations.</li> </ul> <p>24.1.2 Required Readings</p> <p>24.1.2.1 Braithwaite, A. and Smith, F. J., <u>Chromatographic Methods</u>, Chapman and Hall, New York, NY, 1985.</p> <p>24.1.2.2 Moffat, A. C., <u>Clarke's Isolation and Identification of Drugs</u>, The Pharmaceutical Press, London, England, 1986, pp. 160-177.</p> <p>24.1.3 Questions</p> <p>The trainee will provide written answers to the following questions:</p> <ul style="list-style-type: none"> <li>• Define the following: <ul style="list-style-type: none"> <li>○ Thin layer chromatography</li> <li>○ Stationary phase</li> <li>○ Mobile phase</li> <li>○ Solvent front</li> <li>○ R<sub>f</sub> value</li> <li>○ Adsorption</li> <li>○ Absorption</li> <li>○ Elution</li> <li>○ Partition coefficient (K)</li> <li>○ Polarity</li> <li>○ Dipole moment</li> <li>○ Dielectric constant</li> <li>○ Visualizing reagent</li> </ul> </li> <li>• For the silica gel GF TLC plates, what are the “GF” components and what is their purpose?</li> <li>• What is meant by quenching fluorescence?</li> <li>• Why is silica typically chosen over alumina as a stationary phase?</li> <li>• What is the general limit of detection of TLC? What factors influence this?</li> <li>• What are “tailing” and “bearding”? What causes these to occur and what can be done to minimize these effects?</li> <li>• What is an elutropic series? How will the polarity of solvents change when they are mixed together?</li> <li>• Explain the interaction of the sample, mobile phase and stationary phase.</li> <li>• Why do spots with larger R<sub>f</sub> values generally have larger diameters than spots with relatively low R<sub>f</sub> values?</li> <li>• Does sample concentration have an effect on TLC migration? Why or why not?</li> <li>• How can the results of TLC be documented?</li> </ul> <p>24.1.4 Practical Exercises</p>	

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<p>24.1.4.1 The trainee will prepare a TLC bath and visualizing reagent, if applicable, as directed by the trainer.</p> <p>24.1.4.2 The trainee will analyze a set of samples provided by the trainer and will document the TLC results.</p> <p>19.1.5 Evaluation</p> <p>19.1.5.1 The trainer will review the written answers to the questions with the trainee.</p> <p>19.1.5.2 The trainer and the trainee will review and discuss the pertinent points of each of the required readings.</p> <p>19.1.5.3 Review of practical exercises.</p> <p>19.1.5.4 The trainee will be quizzed orally upon the subject matter.</p> <p><b>24.2 Competency Evaluation and Mock Trial</b></p> <p>The trainee may use thin layer chromatography when completing their subdiscipline competency test and may defend their results as a part of their mock trial in that subdiscipline.</p> <p>If using thin layer chromatography in general chemical analysis, the supervisor will document completion of this TLC training section via an e-mail to the trainee.</p> <p><b>24.3 Reading List</b></p> <p>24.3.1 Braithwaite, A. and Smith, F. J., <u>Chromatographic Methods</u>, Chapman and Hall, New York, NY, 1985.</p> <p>24.3.2 Moffat, A. C., <u>Clarke's Isolation and Identification of Drugs</u>, The Pharmaceutical Press, London, England, 1986, pp. 160-177.</p> <p style="text-align: right;">◀End</p>	